Writing Ionic Formulas and Naming Ionic Compounds Notes

Writing Formulas: Hint: _____

Example: _____

Steps	How the example looks during each step
1.	
2.	
3.	

Naming Ionic Compounds

Example: _____

Steps	How the example looks during each step
1.	
2.	
3.	

Another Example: TiO

Anion Endings

- If you see a name that ends in _____, you are dealing with an _____, unless _____, _____, or _____.
- If you see a name that ends in _____ or _____, you are dealing with a _____ (Page 178 or back of periodic table)

How to read/name Polyatomic Ions

- The first "rule" looks at the number of _____ in an ion
  - Think of the _____ ion as being the "base" name
  - The _____ prefix _____ an oxygen.
  - _____ will _____ the oxygens by one.
  - Adding _____ to the –ite version will _____ the number of oxygens by another 1
  - In all situations, the charge is _____ affected.

For example, let's look at the polyatomic ions that involve chlorine

- Cl^-
- ClO^-
- ClO_2^-
- ClO_3^-
- ClO_4^-

How to read/name Polyatomic Ions

- "Rule 2": when the prefix _____ is added to a name, a _____ is added to the ion's formula and its charge is increased by _____
- An Example:
  - Carbonate –
  - Bicarbonate –

How to work with polyatomic ions

- Think of them as a _____ ion
- All the atoms _____ _____ and if you need more than one, you have to use _____.
- For example,

  _____ _____

Practice

Cation	Anion	Name	Formula
		Lithium bromide	
			Na_2O
			PbF_2
Mg^{+2}	**PO_4^{-3}**		
			Cu_2CO_3
		Ammonium sulfate	