

Name: _____ Block: _____

Naming Practice

Write the chemical formula of the following substances

Ionic or covalent? 1. Sodium Carbonate _____

Ionic or covalent? 2. Manganese (VII) nitrate _____

Ionic or covalent? 3. Copper (II) phosphate _____

Ionic or covalent? 4. Sulfurous acid _____

Ionic or covalent? 5. Trisulfur diphosphide _____

Ionic or covalent? 6. Aluminum Carbonite _____

Ionic or covalent? 7. Zinc ^(II) nitrite _____

Ionic or covalent? 8. Lead (II) iodide _____

Ionic or covalent? 9. Phosphoric acid _____

Ionic or covalent? 10. Copper (II) perchlorate _____

Ionic or covalent? 11. Iron (II) sulfate _____

Ionic or covalent? 12. Hydrochloric acid _____

Ionic or covalent? 13. Chloric acid _____

Ionic or covalent? 14. Manganese (IV) oxide _____

Ionic or covalent? 15. Dinitrogen pentoxide _____

Ionic or covalent? 16. Bromine _____

Write the name of the following substances

Ionic or covalent? 1. MgCl_2 _____

Ionic or covalent? 2. BaO _____

Ionic or covalent? 3. PO _____

Ionic or covalent? 4. P_2O _____

Ionic or covalent? 5. HBr _____

Ionic or covalent? 6. Sc_2O_3 _____

Ionic or covalent? 7. H_2S _____

Ionic or covalent? 8. H_2SO_4 _____

Ionic or covalent? 9. CoO _____

Ionic or covalent? 10. FeS _____

Ionic or covalent? 11. HI _____

Ionic or covalent? 12. H_2 _____

Ionic or covalent? 13. NaF _____

Ionic or covalent? 14. AlN _____

Ionic or covalent? 15. B_2S_3 _____

Ionic or covalent? 16. SF_6 _____

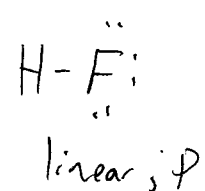
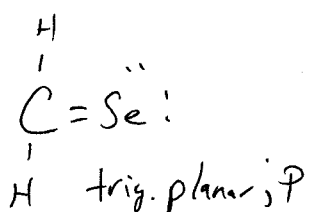
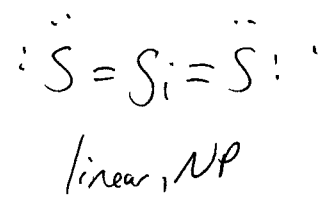
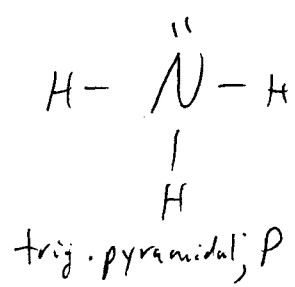
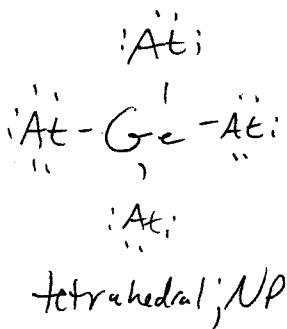
Just to say the red was hard to read.

Name: Key Block: _____

Naming Practice

Write the chemical formula of the following substances

- 1. Sodium Carbonate Na_2CO_3
- 2. Manganese (VII) nitrate $\text{Mn}(\text{NO}_3)_7$
- 3. Copper (II) phosphate $\text{Cu}_3(\text{PO}_4)_2$
- 4. Sulfurous acid H_2SO_3
- 5. Trisulfur diphosphide S_3P_2
- 6. Aluminum Carbonite $\text{Al}_2(\text{CO}_2)_3$ $\text{Al}^{+3}(\text{CO}_2)^{-2}$
- 7. Zinc nitrite $\text{Zn}(\text{NO}_2)_2$ $\text{Zn}^{+2}(\text{NO}_2)^{-1}$
- 8. Lead (II) iodide PbI_2 $\text{Pb}^{+2}(\text{I})^{-1}$
- 9. Phosphoric acid H_3PO_4
- 10. Copper (II) perchlorate $\text{Cu}(\text{ClO}_4)_2$
- 11. Iron (II) sulfate FeSO_4
- 12. Hydrochloric acid HCl
- 13. Chloric acid HClO_3
- 14. Manganese (IV) oxide MnO_2 $\text{Mn}^{+4}\text{O}^{-2}$
- 15. Dinitrogen pentoxide N_2O_5
- 16. Bromine Br_2



Write the name of the following substances

Ionic or covalent? 1. MgCl_2 magnesium chloride

Ionic or covalent? 2. BaO barium oxide

Ionic or covalent? 3. PO phosphorus monoxide

Ionic or covalent? 4. P_2O diphosphorus monoxide

Ionic or covalent? 5. HBr hydrobromic acid

Ionic or covalent? 6. Sc_2O_3 scandium(III) oxide

Ionic or covalent? 7. H_2S hydro sulfic acid or hydrosulfuric acid

Ionic or covalent? 8. H_2SO_4 sulfic acid or sulfuric acid

Ionic or covalent? 9. CoO cobalt(II) oxide

Ionic or covalent? 10. FeS iron(II) sulfide

Ionic or covalent? 11. HI hydro iodic acid

Ionic or covalent? 12. H_2 hydrogen

Ionic or covalent? 13. NaF sodium fluoride

Ionic or covalent? 14. AlN aluminum nitride

Ionic or covalent? 15. B_2S_3 diboron trisulfide

Ionic or covalent? 16. SF_6 sulfur hexafluoride